

Anion Binding in Water Drives Structural Adaptation in an Azaphosphatrane-Functionalized $\text{Fe}^{\text{II}}_4\text{L}_4$ Tetrahedron

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Supporting Information

ABSTRACT: Anion-templated aqueous self-assembly resulted in the formation of an endohedrally functionalized $\text{Fe}^{\text{II}}_4\text{L}_4$ tetrahedron from azaphosphatrane-based subcomponents. This new water-soluble cage is flexible and able to encapsulate anions with volumes ranging from 35 to 219 Å³ via hydrogen bonding and electrostatic interactions. It structurally adapts in response to the size and shape of the template anions, dynamically adopting a conformation either where all four azaphosphatrane ⁺P–H vectors point inward, or else where one points outward and the other three inward. The two cage isomers can coexist in solution and interconvert. A shape memory phenomenon was observed during guest displacement because guest exchange occurs more rapidly than structural reconfiguration.

Self-assembly is an efficient tool for the construction of diverse functional architectures.¹ One subset of these, metal–organic polyhedral complexes, have proven useful in a range of applications, including molecular recognition,² chirality sensing,³ gas separation,⁴ stabilization of reactive species,⁵ and catalysis.⁶ The strategy of subcomponent self-assembly⁷ allows the preparation of a variety of capsules with different shapes and sizes, which bind many different guests.⁸ Most such capsules contain cavities surrounded by aromatic panels.^{7a–d} The elaboration of the cavity microenvironment, through incorporation of functional moieties,^{7e,f} represents an attractive means by which specific substrate encapsulation might be achieved.

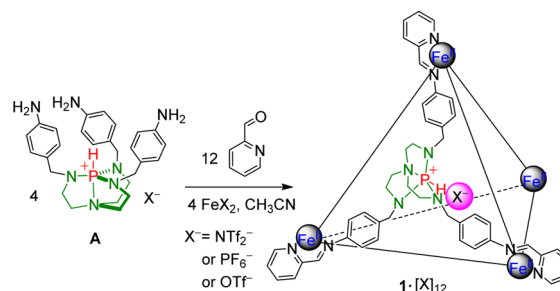
Azaphosphatranes, the conjugate acids of Verkade superbases,⁹ are attractive functional moieties for incorporation into the faces of tetrahedral capsules. Their polarized ⁺P–H groups enable them to be employed as catalysts for CO₂ activation and lactide ring-opening polymerization.¹⁰ Although both of these applications benefit from the hydrogen-bond donor ability of the ⁺P–H group, their use as anion receptors has not been reported. Importantly, these cationic species can be rendered water-soluble via selection of a suitable counterion, potentially enabling their use as anion binders in water, the development of which has been identified as a key challenge in supramolecular chemistry.¹¹

Anion recognition in water is intimately involved in many chemical, biological, and environmental processes.¹¹ However, hydration of both the host and guest often hinders complex-

ation.^{11f} Nonetheless, several synthetic anion receptors able to operate in water have been reported, employing hydrophobic effects, metal-anion coordination, C–H hydrogen bonds, and halogen bonds.^{11f}

Here we report a new azaphosphatrane-functionalized tetrahedron **1** (Scheme 1), assembled from cationic triamine

Scheme 1. Subcomponent Self-Assembly of Cage **1** in Acetonitrile



subcomponent **A**. The design of **1** was based on the following four principles. First, the cationic azaphosphatranes together with the metal ions provide an overall +12 charge, offering a strong electrostatic driving force for anion binding. Second, the cationic framework also inductively increases the affinity of the ⁺P–H hydrogen bond donor for anionic species.¹² Third, the sulfate counterion was chosen to maximize the water solubility of **1**. Fourth, methylene groups between the phenyl rings and azaphosphatrane nitrogen atoms were introduced to enhance cage flexibility, enabling the cavity to fit a diverse array of anions.

We first prepared cage **1** as the bis(trifluoromethanesulfonyl)imide (triflimide, Tf_2N^-), hexafluorophosphate (PF_6^-), and trifluoromethanesulfonate (triflate, TfO^-) salts, all of which were soluble in acetonitrile (Scheme 1). The chloride salt of subcomponent **A** was obtained in four steps (Scheme S1). The Cl^- of **A** readily exchanged with Tf_2N^- , PF_6^- , or TfO^- via anion metathesis. The subsequent reaction of **A** (4 equiv) with the corresponding FeX_2 salt (4 equiv) and 2-formylpyridine (12 equiv) in acetonitrile gave a single product with one azaphosphatrane capping each face (Scheme 1). Products were characterized by NMR and mass spectrometry (Figures S1–S36).

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^{19}F NMR spectra (Figures S5, S17, and S29) indicated the encapsulation of one anion inside the cavity, in slow exchange on the NMR time scale with the external anions.

Single-crystal X-ray diffraction confirmed the structure of **1** in the solid state. As shown in Figure 1, **1** has a face-capped

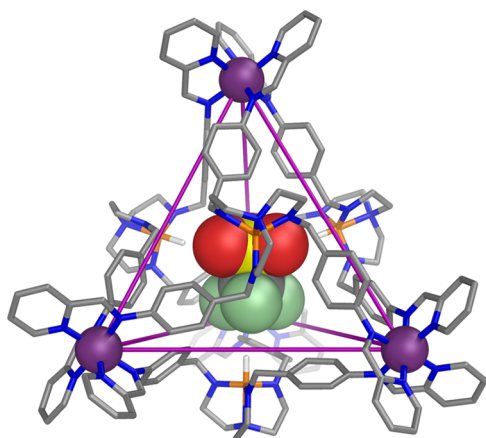


Figure 1. Crystal structure of $\text{TfO}^- \text{C}_1$, which shows a *T*-symmetric cage. Disorder, unbound counterions, non-P-bound hydrogen atoms, and solvents are omitted for clarity.

tetrahedral framework with approximate *T*-symmetry. Both enantiomers ($\Delta\Delta\Delta$ and $\Lambda\Lambda\Lambda$) are present in the unit cell. The $^*\text{P-H}$ groups of the azaphosphatane units point inward into the cavity, where a disordered TfO^- is encapsulated. The Fe^{II} centers are separated by an average of 14.7 \AA , and the volume of the central cavity approximates 166 \AA^3 (51% occupancy for TfO^- , Figure S72).

We then attempted the preparation of a water-soluble cage by employing SO_4^{2-} counterions.¹³ The initial Cl^- of **A** was exchanged for SO_4^{2-} (Scheme S2). However, the subsequent reaction of the sulfate salt of **A** (4 equiv) with FeSO_4 (4 equiv) and 2-formylpyridine (12 equiv) in water failed to give the expected cage complex (Figure S37). We infer that the hydrophilic character of SO_4^{2-} prevents it from acting as a template for the formation of **1**.¹⁴ Therefore, a series of other anions with different shapes and volumes, listed in Figure 2, were added to the aqueous reaction mixture to investigate their template effects (1 equiv of anion was added in each case with regard to the final cage).

All of the anions listed in Figure 2, with volumes ranging from 35 \AA^3 (I^-) to 219 \AA^3 ($\text{CB}_{11}\text{H}_{12}^-$), were observed to serve as competent templates for **1** in water, indicating that its cavity adapted readily to guests having different sizes and shapes.

Cage **1** manifested this adaptability in a novel way: for the smaller template anions, with volumes $\leq 53 \text{ \AA}^3$, part of the population of **1** consisted of a C_3 -symmetric isomer ($\text{C}_3\text{-1}$) in which one of the azaphosphatane $^*\text{P-H}$ groups was oriented away from the inner cavity (*exo*) with the other three pointed inward (*endo*). A *T*-symmetric isomer (*T*-1), containing four *endo* $^*\text{P-H}$ groups, constituted the other part of the population. For the larger anionic templates, having volumes $\geq 55 \text{ \AA}^3$, the *T*-symmetric isomer (*T*-1) of the cage was observed exclusively.

The ^1H NMR spectra of **1** templated by BF_4^- , NO_3^- , or I^- in water thus show five sets of ligand signals (Figures S38) in which one set belongs to *T*-1, while the other four sets in a 1:1:1:1 ratio belong to $\text{C}_3\text{-1}$ (for full structural assignment of $\text{C}_3\text{-1}$, see Figure S43). The corresponding ^{31}P NMR spectra exhibit three peaks, one for *T*-1 and two in a 3:1 ratio for $\text{C}_3\text{-1}$ (Figure S39). The two

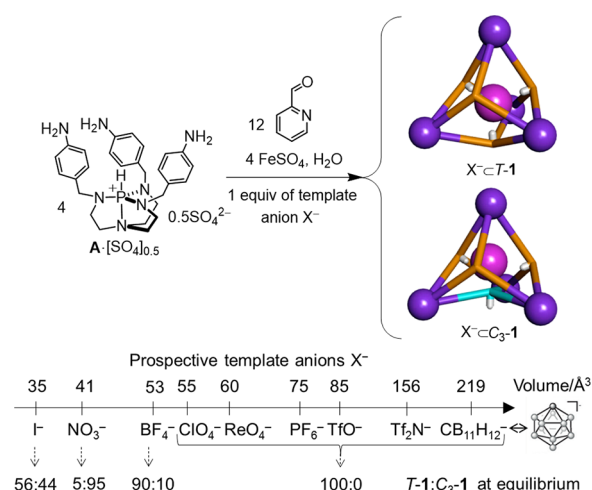


Figure 2. Schematic representation of the two cage isomers $\text{X}^- \text{C-T-1}$ and $\text{X}^- \text{C-C}_3\text{-1}$ obtained by aqueous self-assembly in the presence of 1 equiv of template anion. The cage faces colored orange represent *endo*-azaphosphatane ligands, with $^*\text{P-H}$ pointing inward, while the cyan face represents *exo* $^*\text{P-H}$.

cage isomers were observed to interconvert, taking at least 2 weeks at 298 K to reach equilibrium, as monitored by both ^1H and ^{31}P NMR.

For the self-assembly with NO_3^- as a template, gradual conversion from *T*-1 to $\text{C}_3\text{-1}$ was observed, with the ratio changing from 58:42 at 24 h to 10:90 at 215 h (Figures 3 and S40–

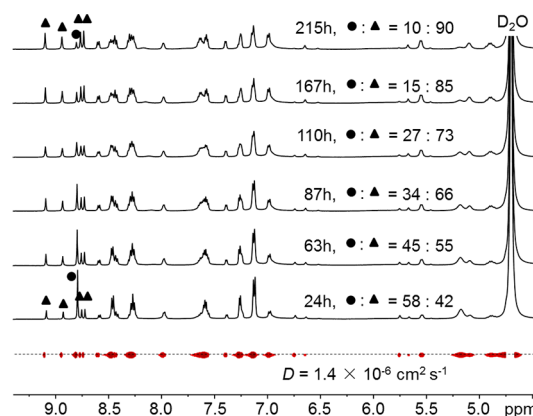


Figure 3. Time evolution of the ratio between $\text{NO}_3^- \text{C-T-1} \cdot [\text{SO}_4]_{5.5}$ (●) and $\text{NO}_3^- \text{C-C}_3\text{-1} \cdot [\text{SO}_4]_{5.5}$ (▲) at 298 K by monitoring the integration of the corresponding imine peaks in the ^1H NMR spectra (D_2O , 500 MHz, 298 K). DOSY NMR indicated that the two species diffuse at the same rate.

S42). This observation indicates that the initially formed mixture between $\text{NO}_3^- \text{C-T-1}$ and $\text{NO}_3^- \text{C-C}_3\text{-1}$ is kinetically metastable. The ratio further shifted to a constant value of 5:95 after 2 weeks at 298 K (Figure S43), corresponding to a Gibbs free energy change (ΔG) of 7.3 kJ mol^{-1} in favor of the C_3 isomer. The free energy of activation (ΔG^\ddagger) from $\text{NO}_3^- \text{C-T-1}$ to $\text{NO}_3^- \text{C-C}_3\text{-1}$ was calculated to be 105 kJ mol^{-1} at 298 K (Supporting Information section 2.4.4 and Figure S68).

In the cases of BF_4^- (Figures S51 and S52) and I^- (Figures S54 and S55), *T*-1 was observed to be thermodynamically favored over $\text{C}_3\text{-1}$, giving equilibrium ratios of 90:10 and 56:44. These values correspond to ΔG differences of 5.4 and 0.64 kJ mol^{-1} , respectively. The ΔG^\ddagger values at 298 K for the conversion from

X^-CT-1 to X^-C_3-1 were 110 and 107 kJ mol⁻¹, for $X^- = BF_4^-$ and I^- , respectively (Figures S69 and S70), similar to the value obtained for NO_3^- . Considering that these kinetic data are close to the previously reported energy barriers for interconversion between analogous diastereomers of M_4L_6 cages (ca. 102 kJ mol⁻¹),¹⁵ and noting that no disassembled products were observed during the conversion, we hypothesize that the face inversion occurs via disruption of $Fe^{II}-N$ linkages, as with M_4L_6 cage diastereomer interconversion,¹⁵ as opposed to a reconfiguration at the azaphosphatranes units, whose P–N bonds are inferred to be stronger.^{9,16} Iminopyridine moieties thus detach from Fe^{II} , allowing the azaphosphatranes to pirouette 180° (Figure S71), whereby the cage framework maintains a degree of structural integrity throughout the rearrangement.¹⁷

Based on the crystal structure of TfO^-C1 in Figure 1, we prepared an MM3 model of $NO_3^-C_3-1$ using the SCIGRESS software package (Figure 4).¹⁸ The trigonal planar NO_3^- is

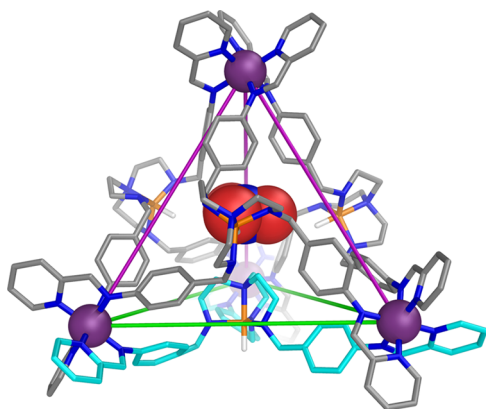


Figure 4. MM3-optimized molecular model of $NO_3^-C_3-1$ based on the crystal structure of TfO^-C1 shown in Figure 1. Non-P-bound hydrogen atoms are omitted for clarity. The *exo*-azaphosphatranes ligand is colored cyan.

centrally bound by the three *endo* $+P-H$ groups. The methylene groups of the *exo*-azaphosphatranes occupy the rest of the cavity. The volume of the preorganized planar cavity approximates 59 Å³ (69% occupancy for NO_3^- , Figure S72), significantly smaller than that of TfO^-CT-1 (166 Å³). This model also suggests that the shape-fit between C_3-1 and NO_3^- may account for the greater favoring by this anion of this isomer, as compared to BF_4^- (Figure S53) or I^- (Figure S56).

¹H, ¹³C, and ³¹P NMR spectra of **1** templated by ClO_4^- , ReO_4^- , PF_6^- , TfO^- , Tf_2N^- , or $CB_{11}H_{12}^-$ in water show only one set of ligand signals (Figures S57–S60), consistent with the exclusive formation of *T-1*. The ¹H–¹⁹F HOESY spectra of the TfO^- and PF_6^- adducts of *T-1* exhibited strong NOE correlations between the F atoms of the encapsulated guest and the $+P-H$ of the azaphosphatranes, underlining the hydrogen bonding interactions (Figures S61, S66, and S67).

Although the cage was unstable in the absence of a template anion, competitive guest exchange offers a way of measuring the relative binding affinities of the different anions. We first chose $TfO^-C1 \cdot [SO_4]_{5.5}$ as the reference and monitored the decrease of the encapsulated TfO^- and increase of the free TfO^- by ¹⁹F NMR upon addition of a competitive anion. Results showed that the binding affinities of PF_6^- , ReO_4^- , and ClO_4^- were 17, 1.9, and 0.028, respectively, relative to that of TfO^- (Figures S73–S75). No obvious TfO^- displacement was observed upon addition of NO_3^- , BF_4^- , I^- , Tf_2N^- , or $CB_{11}H_{12}^-$.¹⁹ We thus chose $Tf_2N^-C1 \cdot$

$[SO_4]_{5.5}$ as the reference to determine the hierarchy of binding affinities of these latter anions. We observed by ¹H NMR that 1.7 equiv of $CB_{11}H_{12}^-$ completely displaced the encapsulated Tf_2N^- (Figure S76), while a larger excess (>6 equiv) of BF_4^- or I^- was needed for Tf_2N^- displacement, with BF_4^- binding more strongly than I^- (Figures S77 and S78). NO_3^- (12 equiv) was unable to displace the bound Tf_2N^- . Combining these results, we observe the following anion binding hierarchy in water: $PF_6^- > ReO_4^- > TfO^- > ClO_4^- > CB_{11}H_{12}^- > Tf_2N^- > BF_4^- > I^- > NO_3^-$.

Although tetrahedron **1** templated by BF_4^- was observed in two distinct isomeric forms ($T-1 \rightleftharpoons C_3-1$), we noted during the titration of BF_4^- into Tf_2N^-CT-1 that only BF_4^-CT-1 was observed within 5 h. $BF_4^-C_3-1$ appeared only thereafter, indicating a short-term shape memory effect (Figure 5). Further

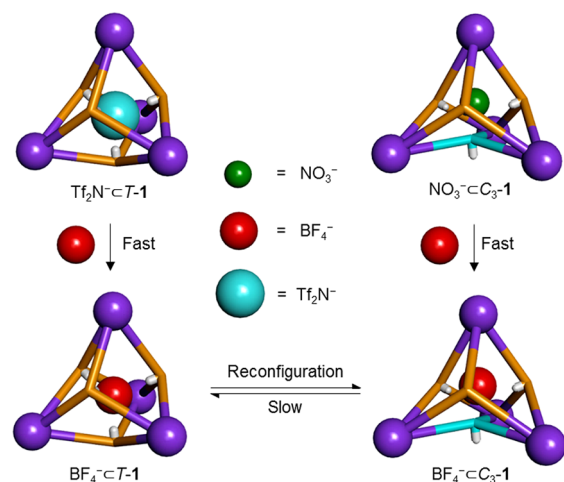


Figure 5. Schematic representation of two successive processes: fast guest displacement of Tf_2N^- or NO_3^- by BF_4^- , giving rise to a shape memory phenomenon, followed by slow isomeric reconfiguration.

experiments showed that, after addition of BF_4^- (2 equiv) to 5:95 ratio of $NO_3^-CT-1:NO_3^-C_3-1$ in solution, a 5:95 ratio of $BF_4^-CT-1:BF_4^-C_3-1$ was observed immediately following the anion displacement (Figure S79). $BF_4^-C_3-1$ then gradually converted to BF_4^-CT-1 . These results indicate the occurrence of two successive processes: rapid guest displacement, with maintenance of the initial ratio of cage diastereomers, followed by slower BF_4^- templated isomeric reconfiguration.

Taking advantage of the larger population of $BF_4^-C_3-1$ (95%) following anion exchange from NO_3^-C1 , we were able to observe heteronuclear NOE correlations between the bound BF_4^- and the methylene protons of the *exo*-azaphosphatranes of $BF_4^-C_3-1$ in ¹H–¹⁹F HOESY spectrum (Figure S80). These correlations provide important structural evidence for the *exo* configuration of one face of C_3-1 , but they could not be observed directly in the BF_4^- -templated assembly, wherein the population of $BF_4^-C_3-1$ constituted less than 25%.

In summary, this study has demonstrated the feasibility of introducing azaphosphatranes into polyhedral complexes using subcomponent self-assembly and also shown the utility of azaphosphatranes as anion binding moieties. The designed flexibility of the structures reported herein enables the aqueous binding of anions with volumes ranging from 35 to 219 Å³ via hydrogen bonding and electrostatic interactions. Structural adaptation of the cage framework was observed in response to the sizes and shapes of template anions, enabling the selective formation of either a *T*- or a *C*₃-symmetric cage framework, which

are capable of interconversion. A shape memory phenomenon was observed during guest displacement, due to its faster kinetics compared to those of framework reconfiguration. Future work will seek to use this tailored cage or one of its derivatives as a catalyst, benefiting from the internally oriented azaphosphatane H-bond donors, for instance, for CO₂ conversion.

■ ASSOCIATED CONTENT

Supporting Information

The Supporting Information is available free of charge on the ACS Publications website at DOI: 10.1021/jacs.7b02950.

Experimental details (PDF)

Crystallographic data for TfO[−]C1 (CCDC 1529220) (CIF)

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Notes

The authors declare no competing financial interest.

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■ REFERENCES

- (1) (a) Wang, W.; Wang, Y. X.; Yang, H. B. *Chem. Soc. Rev.* **2016**, *45*, 2656. (b) Cook, T. R.; Stang, P. J. *Chem. Rev.* **2015**, *115*, 7001. (c) Krieg, E.; Bastings, M. M. C.; Besenius, P.; Rytchinski, B. *Chem. Rev.* **2016**, *116*, 2414. (d) He, Z. F.; Jiang, W.; Schalley, C. A. *Chem. Soc. Rev.* **2015**, *44*, 779. (e) Mauro, M.; Aliprandi, A.; Septiadi, D.; Kehr, N. S.; De Cola, L. *Chem. Soc. Rev.* **2014**, *43*, 4144. (f) Yang, L. L.; Tan, X. X.; Wang, Z. Q.; Zhang, X. *Chem. Rev.* **2015**, *115*, 7196. (g) Yu, G. C.; Jie, K. C.; Huang, F. H. *Chem. Rev.* **2015**, *115*, 7240.
- (2) (a) Chan, A. K. W.; Lam, W. H.; Tanaka, Y.; Wong, K. M. C.; Yam, V. W. W. *Proc. Natl. Acad. Sci. U. S. A.* **2015**, *112*, 690. (b) Bloch, W. M.; Abe, Y.; Holstein, J. J.; Wandtke, C. M.; Dittrich, B.; Clever, G. H. *J. Am. Chem. Soc.* **2016**, *138*, 13750. (c) Zhang, Z.; Kim, D. S.; Lin, C. Y.; Zhang, H. C.; Lammer, A. D.; Lynch, V. M.; Popov, I.; Miljanic, O. S.; Anslyn, E. V.; Sessler, J. L. *J. Am. Chem. Soc.* **2015**, *137*, 7769. (d) Custelcean, R.; Bonnesen, P. V.; Duncan, N. C.; Zhang, X. H.; Watson, L. A.; Van Berkel, G.; Parson, W. B.; Hay, B. P. *J. Am. Chem. Soc.* **2012**, *134*, 8525. (e) Samanta, S. K.; Moncelet, D.; Briken, V.; Isaacs, L. *J. Am. Chem. Soc.* **2016**, *138*, 14488. (f) Bruns, C. J.; Fujita, D.; Hoshino, M.; Sato, S.; Stoddart, J. F.; Fujita, M. *J. Am. Chem. Soc.* **2014**, *136*, 12027.
- (3) (a) You, L.; Berman, J. S.; Anslyn, E. V. *Nat. Chem.* **2011**, *3*, 943. (b) Albrecht, M.; Isaak, E.; Baumert, M.; Gossen, V.; Raabe, G.; Frohlich, R. *Angew. Chem., Int. Ed.* **2011**, *50*, 2850.
- (4) (a) Wu, Y.; Zhou, X. P.; Yang, J. R.; Li, D. *Chem. Commun.* **2013**, *49*, 3413. (b) Chen, T. H.; Popov, I.; Kaveevivitchai, W.; Chuang, Y. C.; Chen, Y. S.; Daugulis, O.; Jacobson, A. J.; Miljanic, O. S. *Nat. Commun.* **2014**, *5*, 5131.
- (5) Ren, H. F.; Huang, Z. H.; Yang, H.; Xu, H. P.; Zhang, X. *ChemPhysChem* **2015**, *16*, 523.
- (6) (a) Yoshizawa, M.; Tamura, M.; Fujita, M. *Science* **2006**, *312*, 251. (b) Cakmak, Y.; Erbas-Cakmak, S.; Leigh, D. A. *J. Am. Chem. Soc.* **2016**, *138*, 1749. (c) Kaphan, D. M.; Levin, M. D.; Bergman, R. G.; Raymond, K. N.; Toste, F. D. *Science* **2015**, *350*, 1235. (d) Omagari, T.; Suzuki, A.; Akita, M.; Yoshizawa, M. *J. Am. Chem. Soc.* **2016**, *138*, 499. (e) Cullen, W.; Misuraca, M. C.; Hunter, C. A.; Williams, N. H.; Ward, M. D. *Nat. Chem.* **2016**, *8*, 231. (f) Howlader, P.; Das, P.; Zangrando, E.; Mukherjee, P. S. *J. Am. Chem. Soc.* **2016**, *138*, 1668.
- (7) (a) Ronson, T. K.; Zarra, S.; Black, S. P.; Nitschke, J. R. *Chem. Commun.* **2013**, *49*, 2476. (b) Zhou, X. P.; Wu, Y.; Li, D. *J. Am. Chem. Soc.* **2013**, *135*, 16062. (c) Sham, K. C.; Yiu, S. M.; Kwong, H. L. *Inorg. Chem.* **2013**, *52*, 5648. (d) Ren, D. H.; Qiu, D.; Pang, C. Y.; Li, Z. J.; Gu, Z. G. *Chem. Commun.* **2015**, *51*, 788. (e) Yi, S.; Brega, V.; Captain, B.; Kaifer, A. E. *Chem. Commun.* **2012**, *48*, 10295. (f) Young, M. C.; Holloway, L. R.; Johnson, A. M.; Hooley, R. J. *Angew. Chem., Int. Ed.* **2014**, *53*, 9832. (g) Bunzen, H.; Nonappa; Kalenius, E.; Hietala, S.; Kolehmainen, E. *Chem. - Eur. J.* **2013**, *19*, 12978. (h) Dömer, J.; Sliotweg, J. C.; Hupka, F.; Lammertsma, K.; Hahn, F. E. *Angew. Chem., Int. Ed.* **2010**, *49*, 6430. (i) Lewing, D.; Koppetz, H.; Hahn, F. E. *Inorg. Chem.* **2015**, *54*, 7653.
- (8) (a) Frischmann, P. D.; Kunz, V.; Würthner, F. *Angew. Chem., Int. Ed.* **2015**, *54*, 7285. (b) Roukala, J.; Zhu, J. F.; Giri, C.; Rissanen, K.; Lantto, P.; Telkki, V. V. *J. Am. Chem. Soc.* **2015**, *137*, 2464. (c) Luo, D.; Zhou, X. P.; Li, D. *Angew. Chem., Int. Ed.* **2015**, *54*, 6190.
- (9) Lensink, C.; Xi, S. K.; Daniels, L. M.; Verkade, J. G. *J. Am. Chem. Soc.* **1989**, *111*, 3478.
- (10) (a) Chatelet, B.; Joucla, L.; Dutasta, J. P.; Martinez, A.; Szeto, K. C.; Dufaud, V. *J. Am. Chem. Soc.* **2013**, *135*, 5348. (b) Zhang, D. W.; Jardel, D.; Peruch, F.; Calin, N.; Dufaud, V.; Dutasta, J. P.; Martinez, A.; Bibal, B. *Eur. J. Org. Chem.* **2016**, *2016*, 1619.
- (11) (a) Busschaert, N.; Caltagirone, C.; Van Rossom, W.; Gale, P. A. *Chem. Rev.* **2015**, *115*, 8038. (b) Ballester, P. *Chem. Soc. Rev.* **2010**, *39*, 3810. (c) Zhou, Y.; Zhang, J. F.; Yoon, J. *Chem. Rev.* **2014**, *114*, 5511. (d) Katayev, E. A.; Kolesnikov, G. V.; Sessler, J. L. *Chem. Soc. Rev.* **2009**, *38*, 1572. (e) Sakkalingam, P.; Shraberg, J.; Rick, S. W.; Gibb, B. C. *J. Am. Chem. Soc.* **2016**, *138*, 48. (f) Langton, M. J.; Serpell, C. J.; Beer, P. D. *Angew. Chem., Int. Ed.* **2016**, *55*, 4629. (g) Kang, S. O.; Llinares, J. M.; Day, V. W.; Bowman-James, K. *Chem. Soc. Rev.* **2010**, *39*, 3980. (h) Jo, J.; Olsz, A.; Chen, C. H.; Lee, D. *J. Am. Chem. Soc.* **2013**, *135*, 3620. (i) Kubik, S. *Chem. Soc. Rev.* **2010**, *39*, 3648. (j) Wang, Q. Q.; Day, V. W.; Bowman-James, K. *Angew. Chem., Int. Ed.* **2012**, *51*, 2119. (k) Fabbri, L.; Poggi, A. *Chem. Soc. Rev.* **2013**, *42*, 1681. (l) Lisbjerg, M.; Valkenier, H.; Jessen, B. M.; Al-Kerdi, H.; Davis, A. P.; Pittelkow, M. *J. Am. Chem. Soc.* **2015**, *137*, 4948.
- (12) Ayme, J. F.; Beves, J. E.; Campbell, C. J.; Gil-Ramirez, G.; Leigh, D. A.; Stephens, A. J. *J. Am. Chem. Soc.* **2015**, *137*, 9812.
- (13) Bolliger, J. L.; Ronson, T. K.; Ogawa, M.; Nitschke, J. R. *J. Am. Chem. Soc.* **2014**, *136*, 14545.
- (14) (a) Custelcean, R. *Chem. Commun.* **2013**, *49*, 2173. (b) Pandurangan, K.; Kitchen, J. A.; Blasco, S.; Boyle, E. M.; Fitzpatrick, B.; Feeney, M.; Kruger, P. E.; Gunnlaugsson, T. *Angew. Chem., Int. Ed.* **2015**, *54*, 4566.
- (15) Meng, W.; Clegg, J. K.; Thoburn, J. D.; Nitschke, J. R. *J. Am. Chem. Soc.* **2011**, *133*, 13652.
- (16) Verkade, J. G. *Acc. Chem. Res.* **1993**, *26*, 483.
- (17) Castilla, A. M.; Ousaka, N.; Bilbeisi, R. A.; Valeri, E.; Ronson, T. K.; Nitschke, J. R. *J. Am. Chem. Soc.* **2013**, *135*, 17999.
- (18) Stewart, J. J. P. *SCIGRESS*, version 2.9.0; Fujitsu Limited: United States, 2009.
- (19) Addition of Tf₂N[−] and CB₁₁H₁₂[−] resulted in partial precipitation of the cage. One equivalent of Tf₂N[−] precipitates about 9% of the cage.